In the Specification:

Amend as follows at page 18:

Example 3.

$$3AgNO_2 + 3AgNO_3 + 4Mo \rightarrow 3Ag_2O + 4MoO_3 + 3N_2 \qquad \text{(VIII)}$$

C1

A comelt of equimolar amounts of $AgNO_2$ and $AgNO_3$ was mixed with a stoichiometric amount of Mo metal in accordance with equation VIII, i.e., 34.1% by weight $AgNO_2$, 37.6% by weight $AgNO_3$, and 28.3% by weight Mo. An autoignition temperature of 131±2°C was determined for the composition using DSC.

~~Example 4.~~

~~$3LiClO_4 + 4Mo \rightarrow 3LiCl + 4MoO_3$ (IX)~~

~~Lithium perchlorate, $LiClO_4$, was mixed with a stoichiometric amount of Mo in accordance with equation IX, i.e., 45.4% by weight $LiClO_4$ and 54.6% by weight Mo. An autoignition temperature of 147±2°C was determined for the composition using DSC.~~

Example 5.

C2

$$2AgNO_3 + 5Mg \rightarrow Ag_2O + 5MgO + N_2 \qquad \text{(X)}$$

$AgNO_3$ was mixed with a stoichiometric amount of magnesium, Mg, metal in accordance with equation X, i.e., 73.7% by weight $AgNO_3$ and 26.3% by weight Mg. An autoignition temperature of 157±2°C was determined for the composition using DSC.

C3

Example 6.

$$KClO_4 + 2AgNO_3 + 9Mg \rightarrow 9MgO + Ag_2O + KCl + N_2 \qquad \text{(XI)}$$

Amend as follows at page 21:

metal in accordance with equation XVI, i.e., 20.5% by weight $NaNO_3$, 41.0% by weight $AgNO_3$ and 38.5% by weight Mo. The composition autoignited at 217±2°C by DSC analysis.

~~Example 12.~~

~~$3CH_6N_4O_3 + 2Mo \rightarrow 2MoO_3 + N_2 + 3CO + 9H_2$ (XVII)~~

~~Guanidine nitrate, $CH_6N_4O_3$, was mixed with a stoichiometric amount of Mo in accordance with equation XVII, i.e., 60.4% by weight $CH_6N_4O_3$ and 39.6% by weight Mo. The composition autoignited at 230±2°C by DSC analysis.~~

~~This is an underoxidized reaction which leaves some products in an incompletely oxidized state. If there is an external source of oxygen the reaction proceeds according to equation XVIII.~~

~~$3CH6N_4O_3 + 2Mo + 6O_2 \rightarrow 2MoO_3 + N_2 + 3CO_2 + 9H_20$ (XVIII)~~

~~This composition points out the utility of using organic nitrates in autoignition reactions.~~

Example 13.

$$CH_6N_4O_3 + 2AgNO_3 + Mo \rightarrow MoO_3 + 3N_2 + CO_2 + 3H_20 + Ag_2O \quad (XIX)$$

A 1:2 ratio of guanidine nitrate to $AgNO_3$ was mixed with a stoichiometric amount of Mo in accordance with equation XIX, i.e., 21.9% by weight $CH_6N_4O_3$, 60.9% $AgNO_3$ and 17.2% by weight Mo. The composition autoignited at 172±2°C (by DSC).

C5
Concld.

This composition is also an example of organic nitrates in autoignition reactions. However, this composition is fully oxidized, and, therefore, requires no external source of oxygen.